

FIG. 1

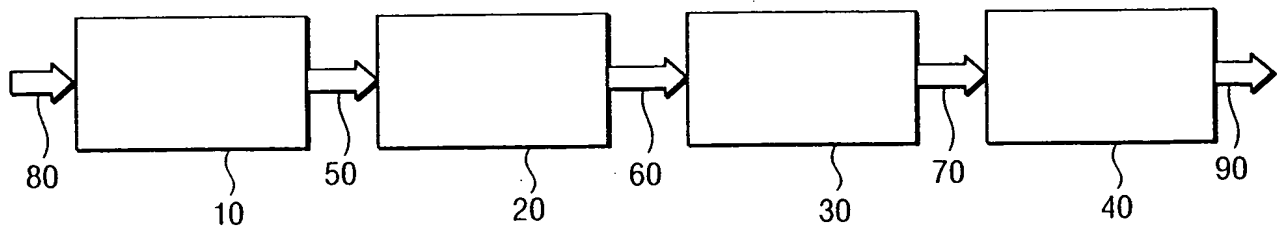


FIG. 2

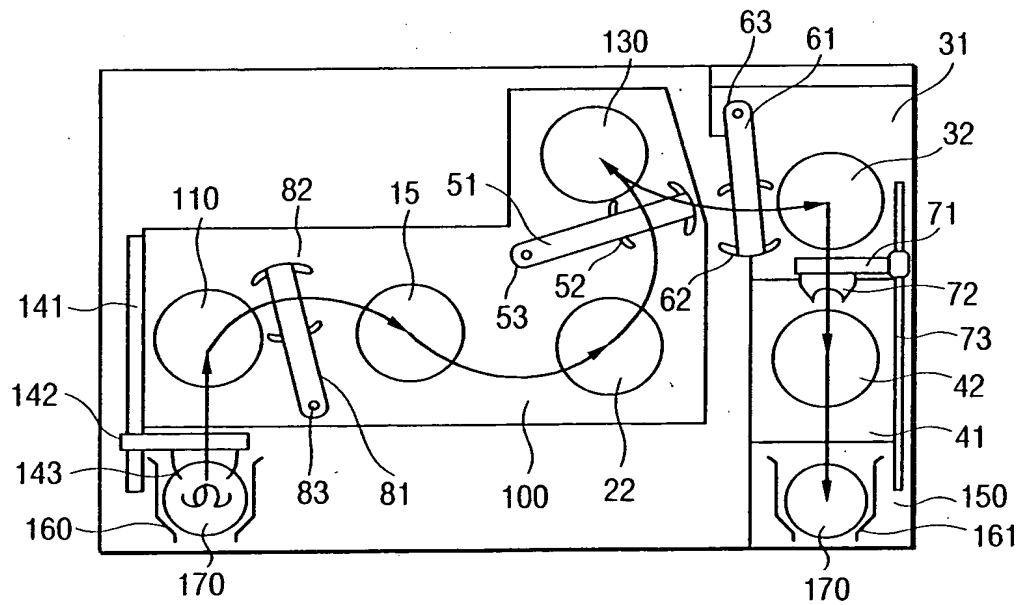


FIG. 3

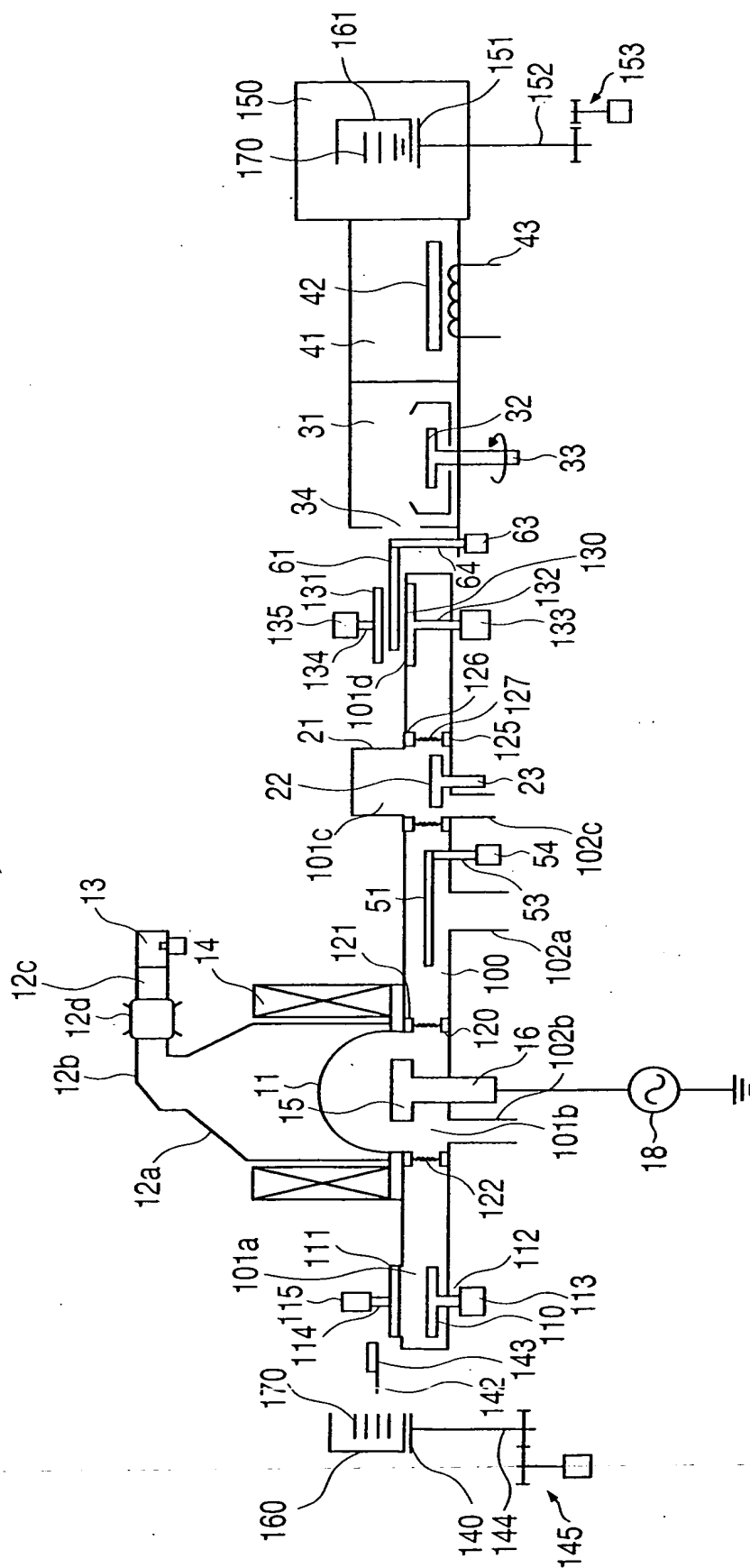


FIG. 4A

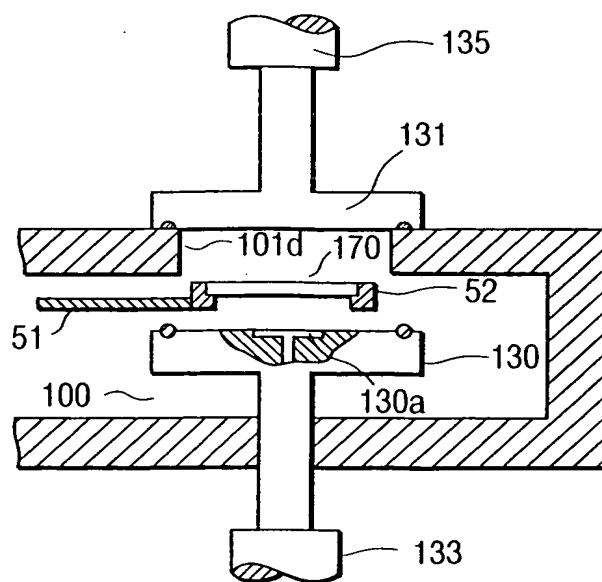


FIG. 4B

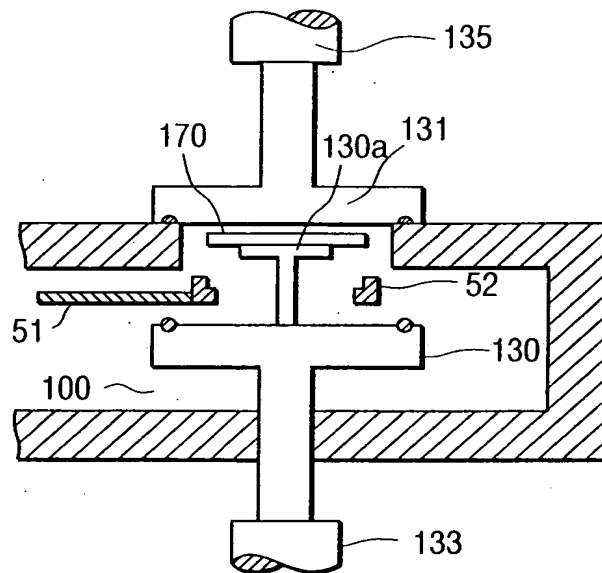


FIG. 4C

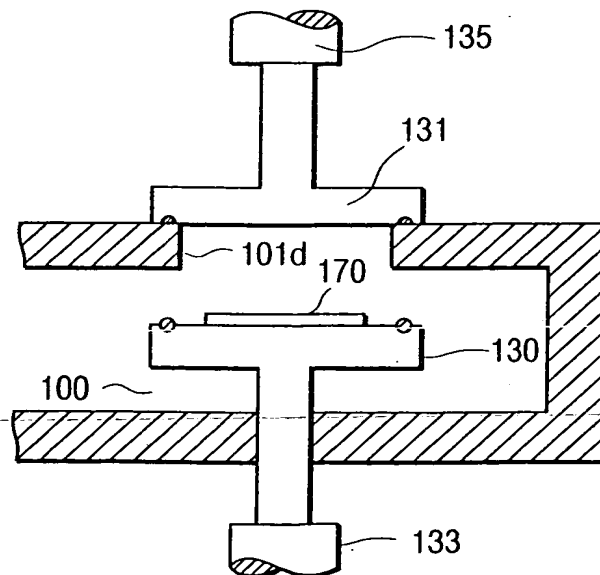


FIG. 4D

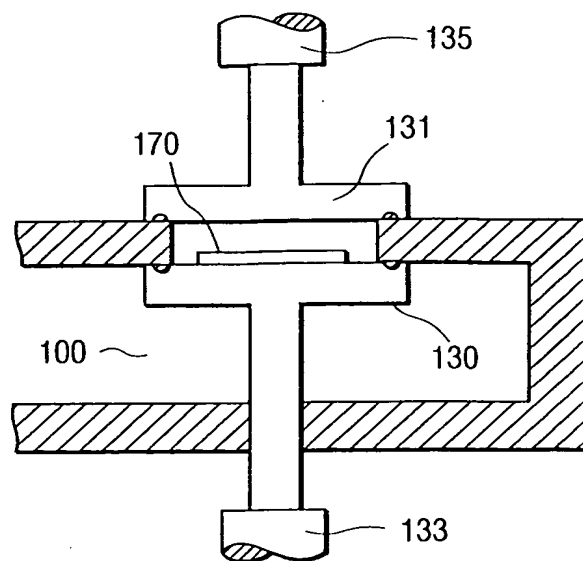


FIG. 4E

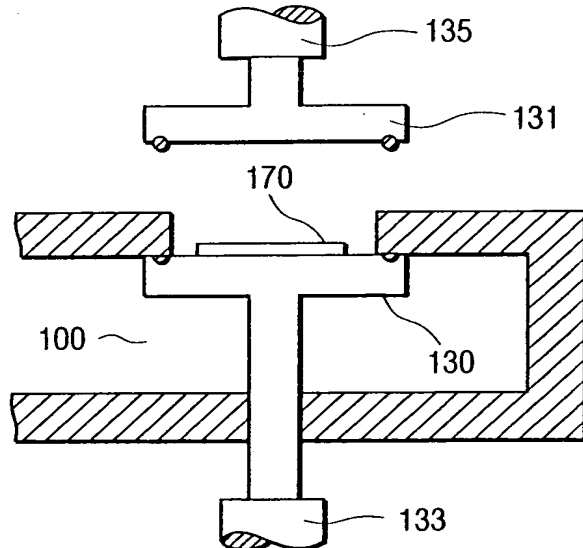


FIG. 4F

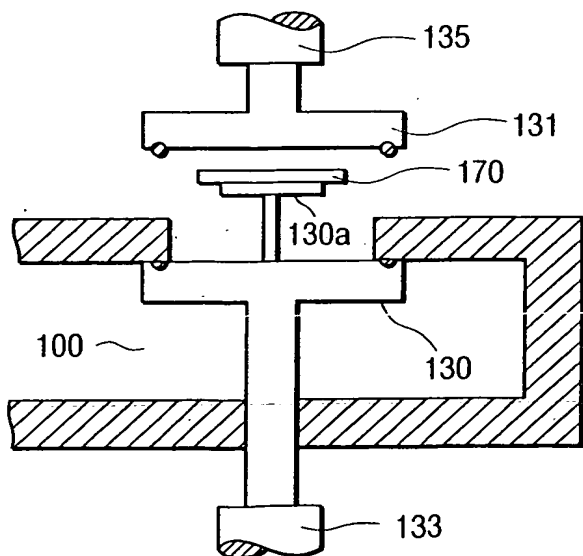


FIG. 4G

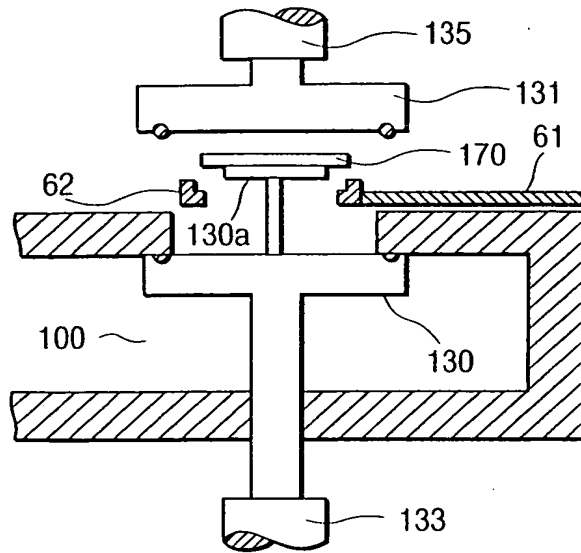


FIG. 5A

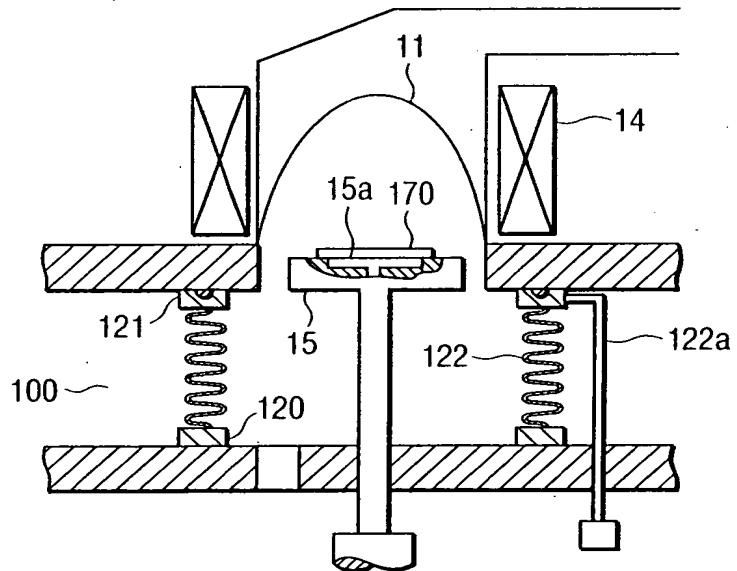


FIG. 5B

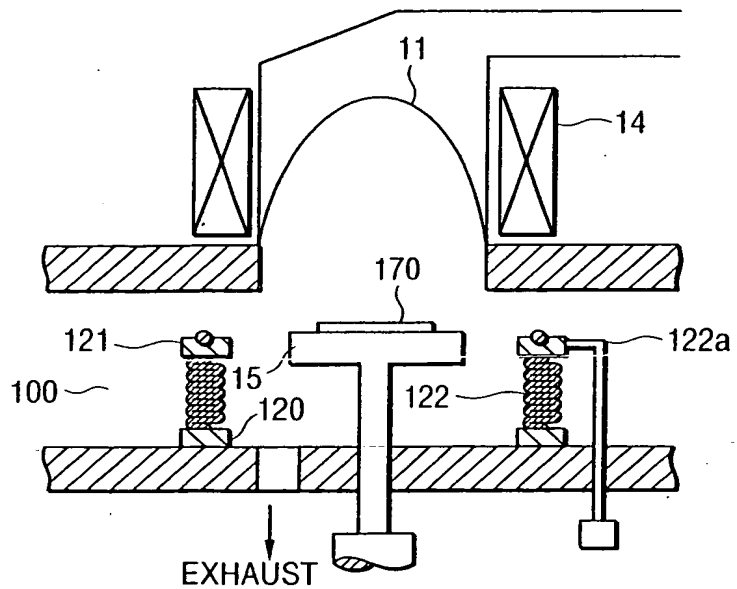


FIG. 6

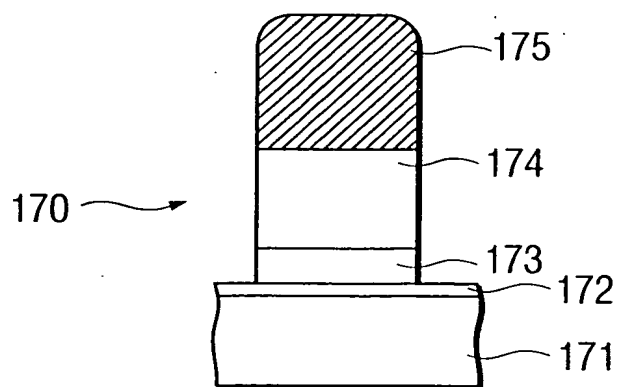
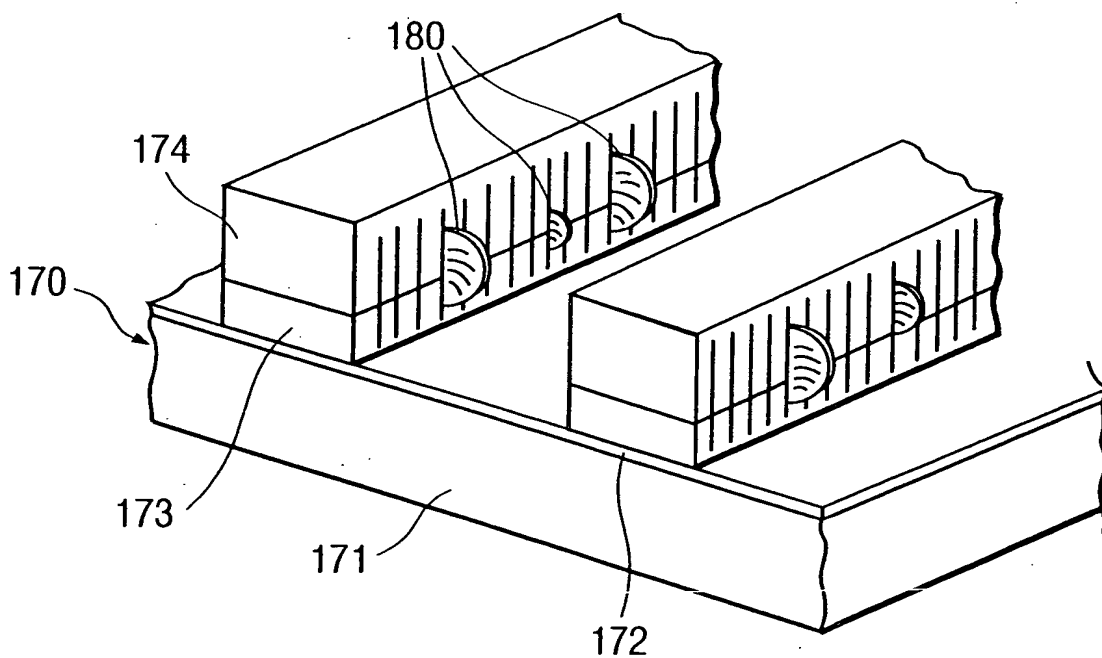


FIG. 7



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FIG. 8

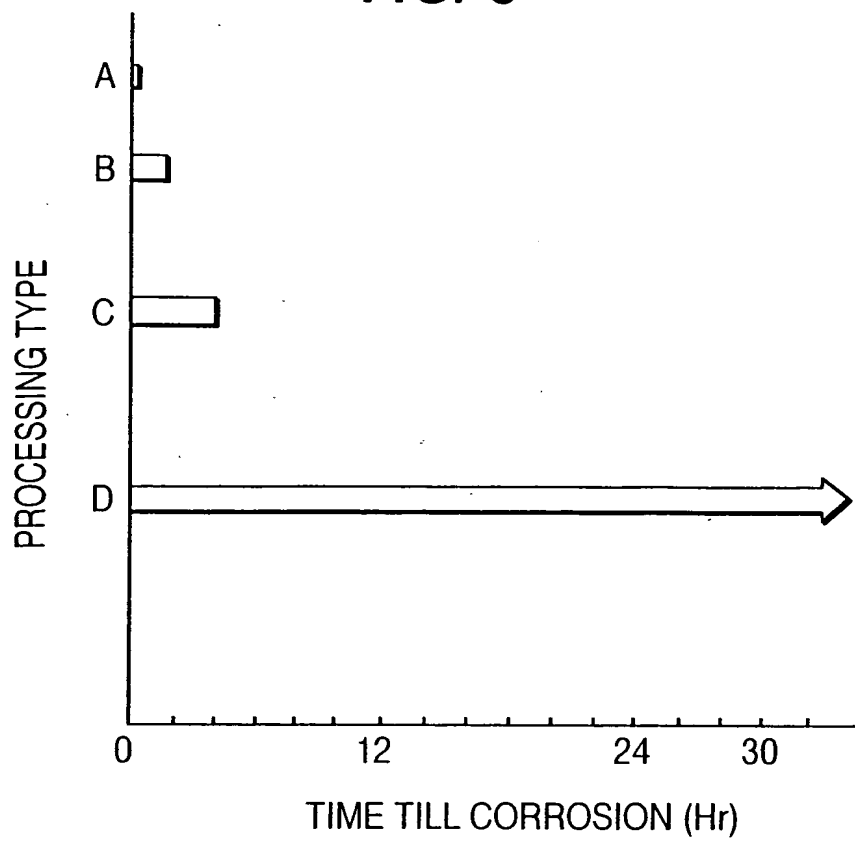


FIG. 9

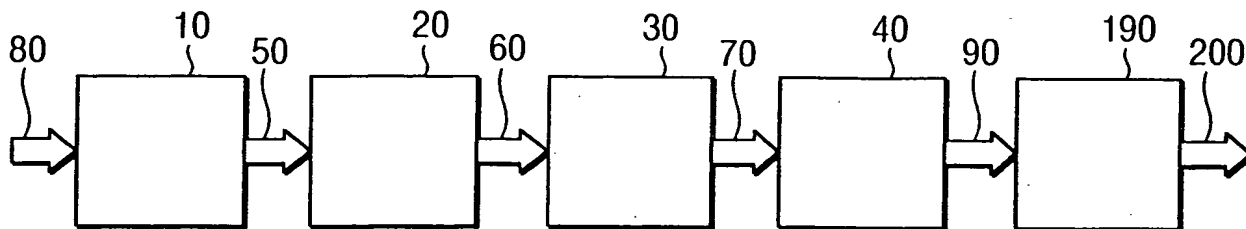


FIG. 10

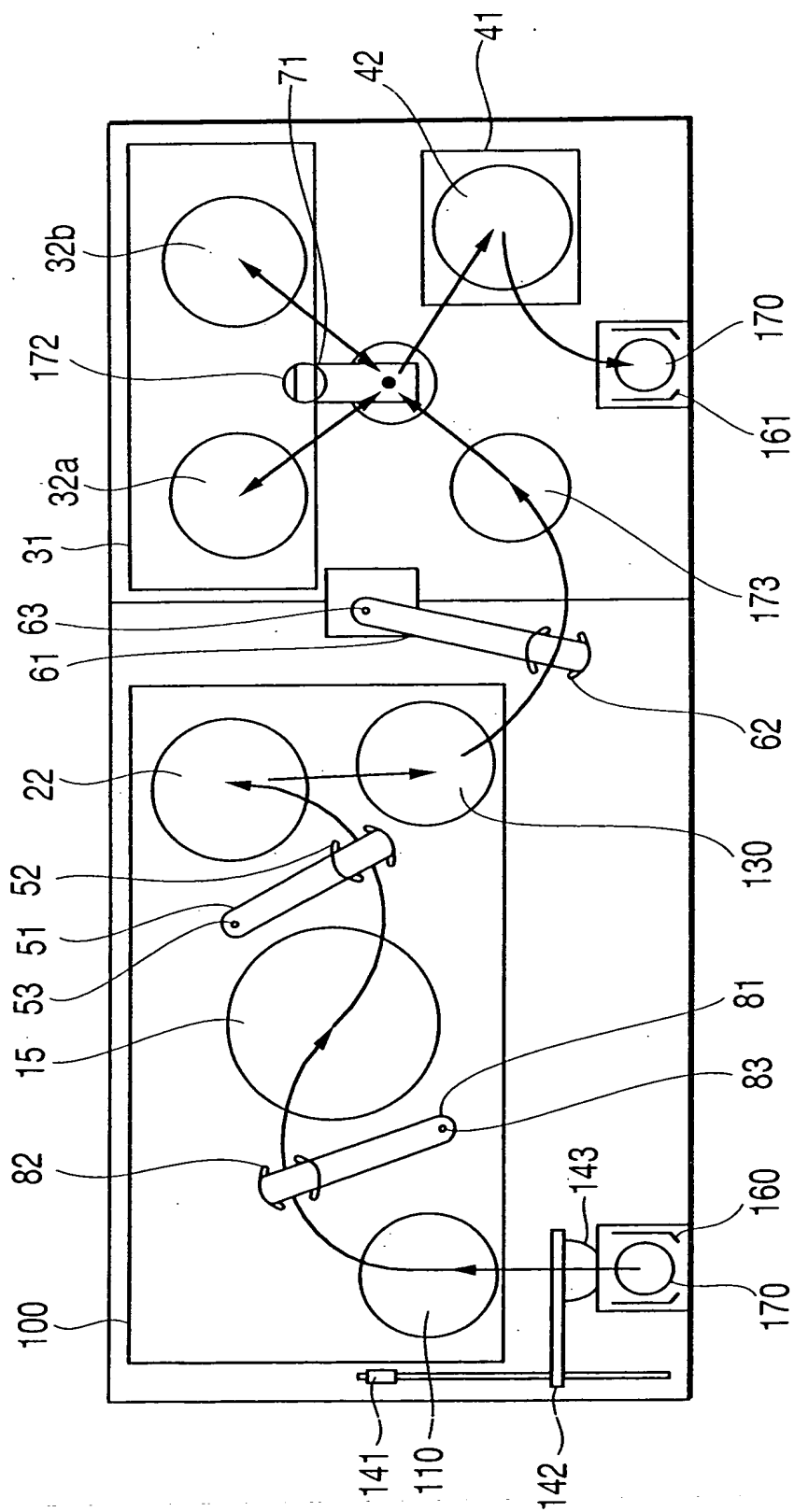


FIG. 11

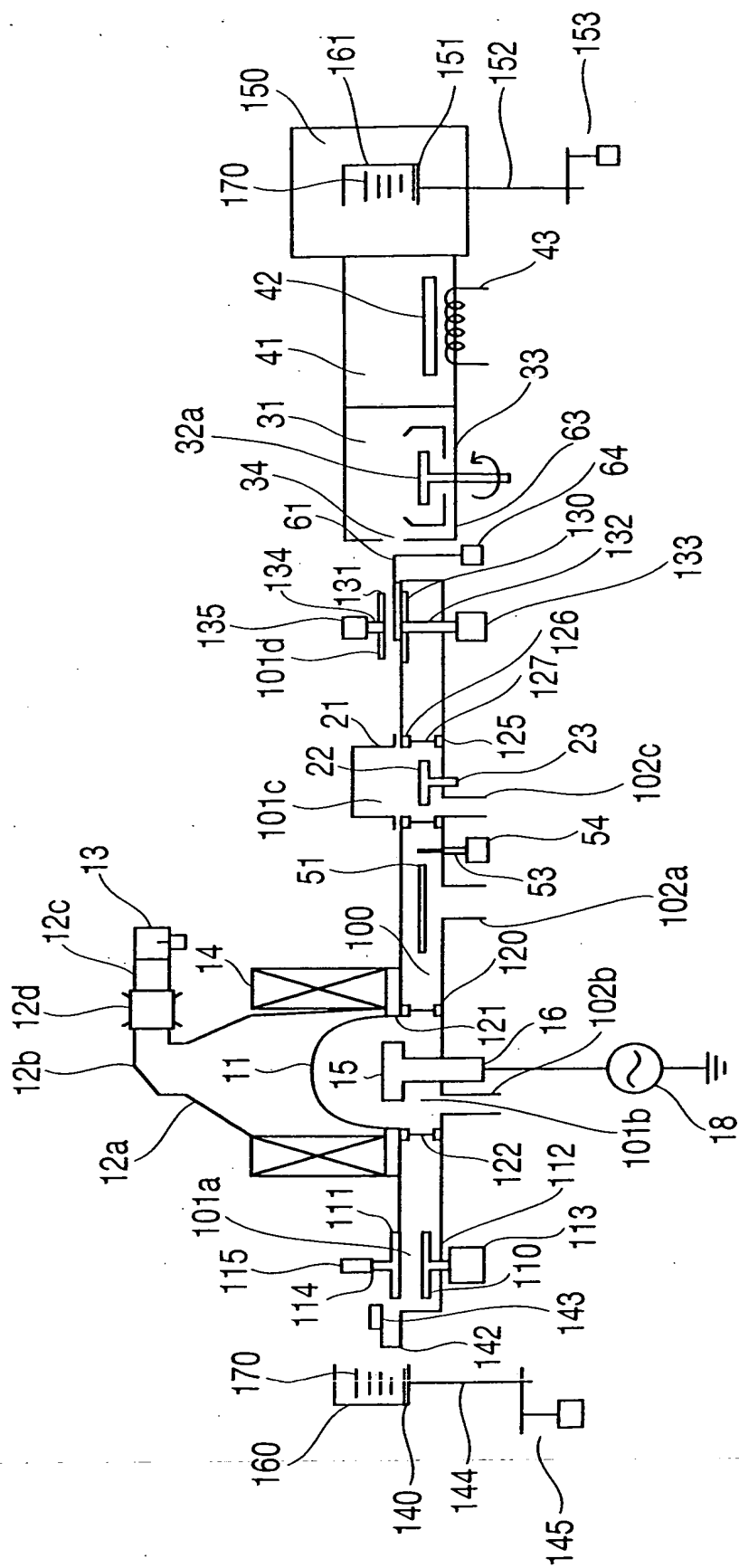
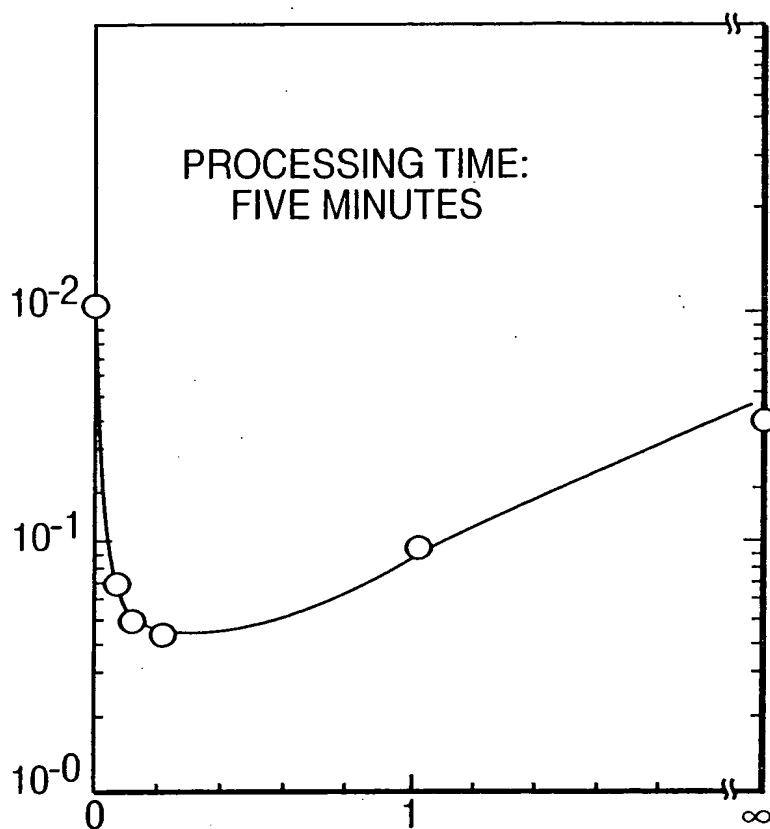


FIG. 12

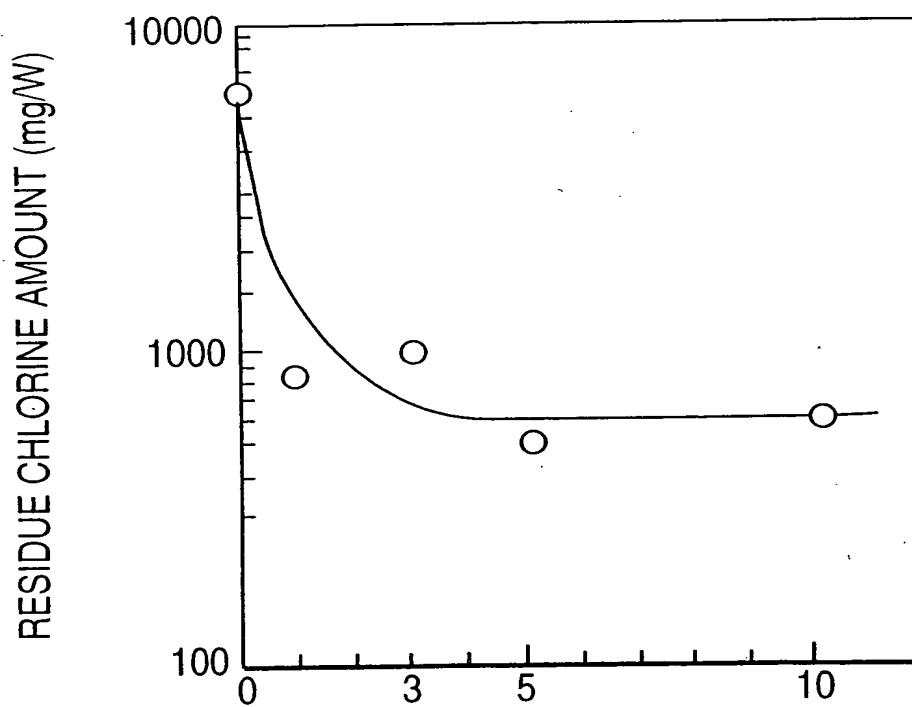
RELATIVE CORROSION OCCURRENCE PERCENTAGE
IN COMPULSORY CORROSION TEST (RELATIVE VALUE)



CH₂COOH/H₂O

ACETIC ACID CONCENTRATION AND
CORROSION OCCURRENCE PERCENTAGE
IN ACETIC ACID WASHING

FIG. 13



PROCESSING (MINUTE)
PROCESSING TIME DEPENDENCY OF RESIDUE CHLORINE

The graph illustrates the titration of acetic acid (CH_3COOH) with ammonium hydroxide (NH_4OH). The x-axis represents the molar ratio of NH_4OH to CH_3COOH , and the y-axis represents the pH of the solution. The curve shows a gradual increase in pH until the ratio reaches 1, where the pH is approximately 5.5. Beyond this point, the pH increases rapidly, reaching approximately 9.5 at a ratio of 2.

$\text{NH}_4\text{OH}/\text{CH}_3\text{COOH}$	pH
0.0	2.5
0.1	3.8
0.2	4.3
0.3	4.6
0.4	4.8
0.5	4.9
0.6	5.0
0.7	5.1
0.8	5.2
0.9	5.3
1.0	5.5
1.1	5.7
1.2	6.5
1.3	8.2
1.4	8.7
1.5	9.0
1.6	9.2
1.7	9.4
1.8	9.5
1.9	9.5
2.0	9.6

FIG. 15(a) ETCHING ON
TABLE 15

A	B	C	D	E	F	G
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FIG. 15(b) ASHING ON
TABLE 22

A	B	C	D	E	F
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FIG. 15(c) RINSING ON
TABLE 32a

	A	C	E
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FIG. 15(d) RINSING ON
TABLE 32b

	B	D
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TIME